

Name: _____

HW5:3 Water the Universal Solvent

Period: _____

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14. What happens to the molecules inside a solvent when it heats up?
15. When temperature increases does more or less gas dissolve in the solvent?
16. When temperature decreases does more or less solid solute dissolve in the solvent?
17. How do cokes keep all of that gas in solution?

LAB QUESTIONS:

18. Even though we didn't have toxic chemicals in the lab, why did we have to wear goggles?

19. When you were trying to understand how temperature affects how fast the sugar dissolves,
 - A) Did you change the way you stirred it?
 - B) Did you change the particle size?
 - C) Did you change the amount of water?
 - D) Did you change the amount of sugar?
 - E) Did you change the temperature of the water?

20. SO, in a good experiment, how many variables do you change?

Get your conclusion statement in!

Assigned: Tues., 12/5 and Wed., 12/6
Due: Thurs., 12/7 and Fri., 12/8

1. What is a polar molecule?
2. Do polar molecules attract polar or non-polar molecules?
3. Does salt dissolve in water?
4. Is salt polar or non-polar?
5. If something is insoluble, is it polar or non-polar?
6. Draw water, being sure to label the positive and negative side.

8. Will a metal be attracted to the oxygen side or hydrogen side?
9. To which side of water will Neon be attracted to?
10. To which side of water will Bromine be attracted to?
11. How could get a solute to dissolve faster in water?

7. Which is the positive side of water: oxygen or hydrogens?

12. Which dissolved faster: granulated sugar or the sugar cube?
 13. Why?
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