| Name:Period:                        |   | HW—25:1L — Acids and Bases<br>Mr. Murray, IPC<br>www.aisd.net/smurray |  | Assigned: Tues., 12/9/03<br>Due: Thurs., 12/9/03 |                 |  |  |
|-------------------------------------|---|---|--|--|-----------------|--|--|
| Together H+ and OH– ions make what? |   | Consider these solutions:   | 5. Which of the following are insoluble in water.                |  |                 |  |  |
|                                     |   | Solution A (pH 4); Solution B (pH 7)<br>Solution C (pH 9)             | MgO  | Li <sub>3</sub> N                                | $S_2F$          |  |  |
|                                     |   | <b>,</b>  | $C_2F_4$   | NaCl   | $NO_2$          |  |  |
|                                     |   | 3. Which is the acid?   | $CO_2$   | Ca2F   | $Na_2O$         |  |  |
| 2 0: 1 4                            | • 1   | 4. Which is neutral?  | _  |  | -               |  |  |
| 2. Circle the acids:                |   | 5. Which is the base?   |  |  |                 |  |  |
| HF                                  | Ca(OH) <sub>2</sub><br>H <sub>2</sub> SO <sub>4</sub> | 6. Which has the most H+ ions?  | Do Vocabulary on the Back  |  |                 |  |  |
| КОН                                 |   | 7. Which has the least H+ ions?                                       |  |  |                 |  |  |
|                                     | 2-14  | 8. Which is equal in H+ and OH ions?                                  |  |  |                 |  |  |
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| Together H+ and OH– ions make what? |   | Consider these solutions:   | <ol><li>Which of the following are insoluble in water.</li></ol> |  |                 |  |  |
|                                     |   | Solution A (pH 4); Solution B (pH 7)                                  | 1110   | ordore in water                                  | •               |  |  |
|                                     |   | Solution C (pH 9)   | MgO  | $Li_3N$  | $S_2F$          |  |  |
|                                     |   | 3. Which is the acid?   | $C_2F_4$   | NaCl   | NO <sub>2</sub> |  |  |
|                                     |   | 4. Which is neutral?  | $CO_2$   | Ca2F   | $Na_2O$         |  |  |
| 2. Circle the acids:                |   | 5. Which is the base?   |  |  |                 |  |  |
| HF Ca(OH) <sub>2</sub>              |   | 6. Which has the most H+ ions?  |  |  |                 |  |  |
|                                     | · /-  | 7. Which has the least H+ ions?                                       | Do Vocabulary on the Back  |  |                 |  |  |
| KOH H <sub>2</sub> SO <sub>4</sub>  |   | 8. Which is equal in H+ and OH ions?                                  |  |  |                 |  |  |
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| Together H+ and OH- ions make what? |   | Consider these solutions:   | 5. Which of the following are insoluble in water.                |  |                 |  |  |
|                                     |   | Solution A (pH 4); Solution B (pH 7)                                  |  |  |                 |  |  |
|                                     |   | Solution C (pH 9)   | MgO  | $Li_3N$  | $S_2F$          |  |  |
|                                     |   | 3. Which is the acid?   | $C_2F_4$   | NaCl   | $NO_2$          |  |  |
|                                     |   | 4. Which is neutral?  | $CO_2$   | Ca2F   | $Na_2O$         |  |  |
| 2. Circle the acids:                |   | 5. Which is the base?   |  |  |                 |  |  |
| HF                                  | Ca(OH) <sub>2</sub>                                   | 6. Which has the most H+ ions?  |  |  |                 |  |  |
|                                     | $H_2SO_4$   | 7. Which has the least H+ ions?                                       | Do V   | Do Vocabulary on the Back                        |                 |  |  |
| КОН                                 |   | 8. Which is equal in H+ and OH ions?                                  |  |  |                 |  |  |
|                                     |   | 1   |  |  |                 |  |  |

|  |   | Vocabulary—<br>don't forget the<br>other side                             | Acid<br>Base<br>Neutral | Neutralize Salt Water Acid Rain Strong Acid pH Weak Base  |  |  |
|--|---|---|-------------------------|---|--|--|
| <ol> <li>1.</li> <li>2.</li> <li>3.</li> <li>4.</li> <li>5.</li> </ol> | A compound that adds a lot of H+ ions because it ionizes completely.  A compound that give OH– ions to water.  When a solution has an equal number of H+ and OH– ions; pH7. |   | 7.<br>8.                | A compound that give just a few OH– ions to water because it does not ionize completely.  Caused by pollution. Has a pH lower than 5.6; is dangerous to living things and buildings.  The product of a neutralization reaction.  To put an equal amount of acid and base together so that their ions cancel each other out. |  |  |
|  |   | Vocabulary—<br>don't forget the<br>other side                             | Acid<br>Base<br>Neutral | Neutralize Salt Water Acid Rain Strong Acid pH Weak Base  |  |  |
| 1.<br>2.<br>3.<br>4.   | completely.  A compound that give OH–   | of H+ ions because it ionizes ions to water. al number of H+ and OH– ions | 7.<br>8.                | A compound that give just a few OH– ions to water because it does not ionize completely.  Caused by pollution. Has a pH lower than 5.6; is dangerous to living things and buildings.  The product of a neutralization reaction.  To put an equal amount of acid and base together so that their ions cancel each other out. |  |  |
|  |   | Vocabulary—<br>don't forget the<br>other side                             | Acid<br>Base<br>Neutral | Neutralize Salt Water Acid Rain Strong Acid pH Weak Base  |  |  |
| 1.<br>2.<br>3.   | The scale to measure acids a A compound that adds a lot completely. A compound that give OH—  | of H+ ions because it ionizes   |                         | A compound that give just a few OH– ions to water because it does not ionize completely.  Caused by pollution. Has a pH lower than 5.6; is dangerous to living things and buildings.  |  |  |

When a solution has an equal number of H+ and OH- ions;

5. A compound that gives H+ ions to water.

pH7.

The product of a neutralization reaction.

9. To put an equal amount of acid and base together so that their ions cancel each other out.