Name: _____

Period:_____

HW—25:1L — Acids and Bases
Mr. Murray, IPC
www.aisd.net/smurray

Assigned: Tues., 12/9/03
Due: Thurs., 12/9/03

1. Together H+ and OH– ions make what?

2. Circle the acids:

HF	$Ca(OH)_2$
KOH	H_2SO_4

Consider these solutions:

Solution A (pH 4); Solution B (pH 7)
Solution C (pH 9)

3. Which is the acid?
4. Which is neutral?
5. Which is the base?
6. Which has the most H+ ions?
7. Which has the least H+ ions?
8. Which is equal in H+ and OH ions?

5. Which of the following are insoluble in water.

MgO	Li_3N	S_2F
C_2F_4	NaCl	NO_2
CO_2	Ca2F	Na_2O

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Vocabulary— don't forget the other side

Acid	Neutralize	Salt Water
Base	Acid Rain	Strong Acid
Neutral	pH	Weak Base

1. The scale to measure acids and bases.
2. A compound that adds a lot of H+ ions because it ionizes completely.
3. A compound that give OH– ions to water.
4. When a solution has an equal number of H+ and OH– ions; pH7.
5. A compound that gives H+ ions to water.
6. A compound that give just a few OH– ions to water because it does not ionize completely.
7. Caused by pollution. Has a pH lower than 5.6; is dangerous to living things and buildings.
8. The product of a neutralization reaction.
9. To put an equal amount of acid and base together so that their ions cancel each other out.

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