Name: _____

Period:_____

HW-25:1 - Acids and Bases Mr. Murray, IPC www.aisd.net/smurray

Assigned: Tues.,	12/9/03
Due: Thurs.,	12/9/03

1. Circle the acids:		Consider these solutions:		
NaOH	H_2SO_3	Solution A (pH 11); Solution B (pH 3) Solution C (pH 7)		
HCl	Mg(OH) ₂	3. Which is the acid?		
		4 Which is a system 19		

2. Together H+ and OH- ions make what?

Consider these solutions:				
Solution A (pH 11); Solution B (pH 3) Solution C (pH 7)				
3. Which is the acid?				
4. Which is neutral?				
5. Which is the base?				
6. Which has the most H+ ions?				
7. Which has the least H+ ions?				

8. Which is equal in H+ and OH ions?

5. Which of the following are soluble in water.

MgO	Li ₃ N	S_2F
C_2F_4	NaCl	NO_2
CO_2	Ca2F	Na ₂ O

Do Vocabulary on the Back

Name: Period: 1. Circle the acids:		HW—25:1 — Acids and Bases Mr. Murray, IPC www.aisd.net/smurray		Assigned: Tues., 12/9/03 Due: Thurs., 12/9/03	
		Consider these solutions:	5. Which of the following are soluble in water.		
NaOH	H_2SO_3	Solution A (pH 11); Solution B (pH 3) Solution C (pH 7)	MgO	Li ₃ N	S_2F
HCl	Mg(OH) ₂	3. Which is the acid?	C_2F_4	NaCl	NO ₂
		4. Which is neutral?	CO_2	Ca2F	Na ₂ O
2. Together H+ and OH– ions make what?		5. Which is the base?	Do Vocabulary on the Back		
		6. Which has the most H+ ions?			n the Deels
		7. Which has the least H+ ions?			in the dack
		8. Which is equal in H+ and OH ions?			

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NaOH	H_2SO_3	Solution A (pH 11); Solution B (pH 3) Solution C (pH 7)	MgO	Li ₃ N	S_2F		
HCl	Mg(OH) ₂	3. Which is the acid?	C_2F_4	NaCl	NO ₂		
		4. Which is neutral?	CO_2	Ca2F	Na ₂ O		
2 T (1		5. Which is the base?					
2. Together make what?	H+ and $OH-$ lons	6. Which has the most H+ ions?	Do Vocabulary on the Back				
		7. Which has the least H+ ions?					

8. Which is equal in H+ and OH ions?

Name: Period:	Vocabulary— don't forget the	Acid Base Neutral	Neutralize Acid Rain pH	Salt Water Strong Acid Weak Base
	other side	Neutral	рп	weak base

- 1. When a solution has an equal number of H+ and OH– ions; pH7.
- 2. The scale to measure acids and bases.
- 3. To put an equal amount of acid and base together so that their 8. ions cancel each other out.
- 4. The product of a neutralization reaction.
- 5. A compound that adds a lot of H+ ions because it ionizes completely.

- 6. Caused by pollution. Has a pH lower than 5.6; is dangerous to living things and buildings.
- 7. A compound that give OH– ions to water.
- 8. A compound that gives H+ ions to water.
- 9. A compound that give just a few OH– ions to water because it does not ionize completely..

Name:		Vocabularv—	Acid		Neutralize	Salt Water
Period:_		don't forget the B other side N			Acid Rain pH	Strong Acid Weak Base
1.	When a solution has an equa pH7.	l number of H+ and OH– ions	; 6.	Caused by pol to living thing	lution. Has a pH lower s and buildings.	than 5.6; is dangerous
2.	 The scale to measure acids and bases. A compound that give OH- ions to water. 		ater.			
3.	To put an equal amount of acions cancel each other out.	cid and base together so that th	neir 8.	8. A compound that gives H+ ions to water.		iter.
4.	The product of a neutralizati	on reaction.	9.	 A compound that give just a few OH- ions to water bec it does not ionize completely 		
5.	A compound that adds a lot o completely.	of H+ ions because it ionizes				

Name: _		Vocabularv—	Acid		Neutraliza	Salt Water
Period:_		don't forget the other side	Base Neutral		Acid Rain pH	Strong Acid Weak Base
1.	When a solution has an equal pH7.	l number of H+ and OH– ions;	; 6.	Caused by pol to living thing	llution. Has a pH lower s and buildings.	than 5.6; is dangerous
2.	2. The scale to measure acids and bases.		7.	A compound that give OH- ions to water.		ater.
3.	To put an equal amount of actions cancel each other out.	eid and base together so that th	eir 8.	8. A compound that gives H+ ions to water.		iter.
4.	The product of a neutralization	on reaction.	9.	 A compound that give just a few OH- ions to water beca it does not ionize completely 		
5.	A compound that adds a lot o completely.	of H+ ions because it ionizes				