

Name: _____

Period: _____

HW—25:1 — Acids and Bases
Mr. Murray, IPC
www.aisd.net/smurray

Assigned: Tues., 12/9/03
Due: Thurs., 12/9/03

1. Circle the acids:

NaOH H₂SO₃
HCl Mg(OH)₂

2. Together H⁺ and OH⁻ ions make what?

Consider these solutions:

Solution A (pH 11); Solution B (pH 3)
Solution C (pH 7)

3. Which is the acid?
4. Which is neutral?
5. Which is the base?
6. Which has the most H⁺ ions?
7. Which has the least H⁺ ions?
8. Which is equal in H⁺ and OH ions?

5. Which of the following are soluble in water.

MgO Li₃N S₂F
C₂F₄ NaCl NO₂
CO₂ Ca₂F Na₂O

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**Vocabulary—
don't forget the
other side**

Acid
Base
Neutral

Neutralize
Acid Rain
pH

Salt Water
Strong Acid
Weak Base

1. When a solution has an equal number of H^+ and OH^- ions; pH7.
2. The scale to measure acids and bases.
3. To put an equal amount of acid and base together so that their ions cancel each other out.
4. The product of a neutralization reaction.
5. A compound that adds a lot of H^+ ions because it ionizes completely.
6. Caused by pollution. Has a pH lower than 5.6; is dangerous to living things and buildings.
7. A compound that give OH^- ions to water.
8. A compound that gives H^+ ions to water.
9. A compound that give just a few OH^- ions to water because it does not ionize completely..

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