

Name: _____

Period: _____

HW—21:Rev —Review for Test
Mr. Murray, IPC
www.aisd.net/smurray

Assigned: Thurs., 11/20/03
Due: Mon., 11/24/03

Type of Reaction	Balance the reactions:
1. _____	____ FeCl ₃ + ____ Na(OH) → ____ Fe(OH) ₃ + ____ NaCl
2. _____	____ KClO ₃ → ____ KCl + ____ O ₂
3. _____	____ C ₃ H ₈ + ____ O ₂ → ____ CO ₂ + ____ H ₂ O
4. _____	____ Na + ____ H ₂ O → ____ Na(OH) + ____ H ₂
5. _____	____ S ₈ + ____ O ₂ → ____ SO ₃

6. Is combustion endothermic or exothermic?

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**Vocabulary
chapter 20 and 21**

Closed System
Open System
Products
Reactants

Limiting Reactant
Endothermic
Exothermic

Law of Conservation
of Mass
Principle of Definite
Proportions

1. An open beaker would be an example of this:
2. The reactant that is used up first and limits the reaction.
3. A flask with a balloon on it is an example of this:
4. In a chemical reaction the arrow points to this:
5. The left side of a chemical reaction; the arrow points from here.

6. Says that the reactants must equal the products (chemistry is science not magic).
7. A reaction that gets hot, like combustion (produces energy).
8. A reaction that gets cold (absorbs energy).
9. Says that compounds have to be only one way (H_2O is water, but not HO or H_3O).

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