#### HW—21:Rev —Review for Test Mr. Murray, IPC www.aisd.net/smurray

## Assigned: Thurs., 11/20/03 Due: Mon., 11/24/03

Type of Reaction	Balance the reactions:	
1	$\underline{\qquad} \operatorname{FeCl}_3 + \underline{\qquad} \operatorname{Na(OH)} \rightarrow \underline{\qquad} \operatorname{Fe(OH)}_3 + \underline{\qquad} \operatorname{NaCl}$	
2	$\underline{\qquad} KClO_3 \rightarrow \underline{\qquad} KCl + \underline{\qquad} O_2$	
3	$\underline{\qquad} C_3H_8 + \underline{\qquad} O_2 \rightarrow \underline{\qquad} CO_2 + \underline{\qquad} H_2O$	
4	$\underline{\qquad} Na + \underline{\qquad} H_2O \rightarrow \underline{\qquad} Na(OH) + \underline{\qquad} H_2$	
5	$\underline{\qquad} S_8 + \underline{\qquad} O_2 \rightarrow \underline{\qquad} SO_3$	

6. Is combustion endothermic or exothermic?

Do Vocabulary on the Back

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Type of Reaction	<b>Balance the reactions:</b>	6. Is combustion endothermic or exothermic?
1	$\underline{\qquad} FeCl_3 + \underline{\qquad} Na(OH) \rightarrow \underline{\qquad} Fe(OH)_3 + \underline{\qquad} NaCl$	
2	$\underline{\qquad} \text{KClO}_3 \rightarrow \underline{\qquad} \text{KCl} + \underline{\qquad} \text{O}_2$	
3	$\underline{\qquad} C_3H_8 + \underline{\qquad} O_2 \rightarrow \underline{\qquad} CO_2 + \underline{\qquad} H_2O$	Do Vocabulary on the B
4	$\underline{\qquad} Na + \underline{\qquad} H_2O \rightarrow \underline{\qquad} Na(OH) + \underline{\qquad} H_2$	
5	$\underline{\qquad} S_8 + \underline{\qquad} O_2 \rightarrow \underline{\qquad} SO_3$	

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6. Is combustion endothermic or exothermic?

#### Do Vocabulary on the Back

Name: \_\_\_\_\_

Period:

Name: Period:	Vocabulary chapter 20 and 21	Closed System Open System Products Reactants	Limiting Reactant Endothermic Exothermic
1. An open beaker would be an example of this:		6. Says that th	e reacts must equal the

- 2. The reactant the is used up first and limits the reaction.
- 3. A flask with a balloon on it is an example of this:
- 4. In a chemical reaction the arrow points to this:
- 5. The left side of a chemical reaction; the arrow points from here.

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5. The left side of a chemical reaction; the arrow points

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5. Says that the reacts must equal the products (chemistry is science not magic).

Law of Conservation

of Mass Principle of Definite Proportions

- 7. A reaction that gets hot, like combustion (produces energy).
- 8. A reaction that gets cold (absorbs energy).
- 9. Says that compounds have to be only one way ( $H_2O$  is water, but not HO or  $H_3O$ ).

Name: Period:	Vocabulary chapter 20 and 21	Closed System Open System Products Reactants	Limiting Reactant Endothermic Exothermic	Law of Conservation of Mass Principle of Definite Proportions
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Name:	Vocabularv	Closed System	Limiting Reactant Endothermic Exothermic	Law of Conservation of Mass Principle of Definite Proportions
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9. Says that compounds have to be only one way (H<sub>2</sub>O is water, but not HO or H<sub>3</sub>O).