

Name: \_\_\_\_\_

Period: \_\_\_\_\_

**HW—21:1L—Types of Reactions**  
**Mr. Murray, IPC**  
[www.aisd.net/smurray](http://www.aisd.net/smurray)

**Assigned: Fri., 11/14/03**  
**Due: Tues., 11/18/03**

Type of Reaction	Balance the reactions:
1. _____	____ Cl <sub>2</sub> + ____ Zn <sub>2</sub> O → ____ ZnCl + ____ O <sub>2</sub>
2. _____	____ CH <sub>4</sub> + ____ O <sub>2</sub> → ____ CO <sub>2</sub> + ____ H <sub>2</sub> O
3. _____	____ Na + ____ Cl <sub>2</sub> → ____ NaCl
4. _____	____ Na <sub>2</sub> SO <sub>4</sub> + ____ BaCl <sub>2</sub> → ____ BaSO <sub>4</sub> + ____ NaCl
5. _____	____ Al <sub>2</sub> O <sub>3</sub> + energy → ____ Al + ____ O <sub>2</sub>

6. If a reaction gets hot is it endothermic or exothermic?

7. On the back write a simple addition reaction. Make sure to balance any compounds (if Ionic) and balance the reaction.

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