Name: \_\_\_\_\_

Period:\_\_\_\_\_

**Type of Reaction** 

1. \_\_\_\_\_

2.

3.

4. \_\_\_\_\_

5. \_\_\_\_\_

## HW—21:AT — After Test Mr. Murray, IPC www.aisd.net/smurray

**Balance the reactions:** 

 $\_$   $S_8 + \_$   $O_2 \rightarrow \_$   $SO_3$ 

 $\_$  C<sub>3</sub>H<sub>8</sub> +  $\_$  O<sub>2</sub>  $\rightarrow$   $\_$  CO<sub>2</sub> +  $\_$  H<sub>2</sub>O

 $\underline{\qquad} Na + \underline{\qquad} H_2O \rightarrow \underline{\qquad} Na(OH) + \underline{\qquad} H_2$ 

 $\underline{\qquad} KClO_3 \rightarrow \underline{\qquad} KCl + \underline{\qquad} O_2$ 

 $FeCl_3 +$  Na(OH)  $\rightarrow$  Fe(OH)<sub>3</sub> + NaCl

## Assigned: Thurs., 11/20/03 Due: Mon., 11/24/03

6. You mix vinegar and baking soda. It gets cold— endothermic or exothermic?

Do Vocabulary on the Back

Name: \_\_\_\_\_\_ Period:\_\_\_\_\_\_

HW—21:AT — After Test Mr. Murray, IPC www.aisd.net/smurray

Type of Reaction	Balance the reactions:		
1	$\underline{\qquad} S_8 + \underline{\qquad} O_2 \rightarrow \underline{\qquad} SO_3$		
2	$\underline{\qquad} C_3H_8 + \underline{\qquad} O_2 \rightarrow \underline{\qquad} CO_2 + \underline{\qquad} H_2O$		
3	$\underline{\qquad} Na + \underline{\qquad} H_2O \rightarrow \underline{\qquad} Na(OH) + \underline{\qquad} H_2$		
4	$\underline{\qquad} KClO_3 \rightarrow \underline{\qquad} KCl + \underline{\qquad} O_2$		
5	$\underline{\qquad} FeCl_3 + \underline{\qquad} Na(OH) \rightarrow \underline{\qquad} Fe(OH)_3 + \underline{\qquad} NaCl$		

Assigned: Thurs., 11/20/03 Due: Mon., 11/24/03

6. You mix vinegar and baking soda. It gets cold— endothermic or exothermic?

**Do Vocabulary on the Back** 

Name: \_\_\_\_\_\_ Period: \_\_\_\_\_ HW—21:AT — After Test Mr. Murray, IPC www.aisd.net/smurray

Type of ReactionBalance the reactions:1. $\__S_8 + \__O_2 \rightarrow \__SO_3$ 2. $\__C_3H_8 + \__O_2 \rightarrow \__CO_2 + \__H_2O$ 3. $\__Na + \__H_2O \rightarrow \__Na(OH) + \__H_2$ 4. $\__KClO_3 \rightarrow \__KCl + \__O_2$ 5. $\__FeCl_3 + \__Na(OH) \rightarrow \__Fe(OH)_3 + \__NaCl$ 

## Assigned: Thurs., 11/20/03 Due: Mon., 11/24/03

6. You mix vinegar and baking soda. It gets cold— endothermic or exothermic?

## Do Vocabulary on the Back

Name:		Closed System	Limiting Reactant	Law of Conservation	
Period:		Open System Products Reactants	Endothermic Exothermic	of Mass Principle of Definite Proportions	
1. Says that compounds has but not HO or $H_3O$ ).		6. Says that the reacts must equal the products (chemistry is science not magic).			
2. The reactant the is used	7. A reaction the	7. A reaction that gets hot, like combustion (produces energy).			
3. A reaction that gets cold	d (absorbs energy).	8. An open beaker would be an example of this:			
4. The left side of a chemic from here.	cal reaction; the arrow points	9. A flask with	9. A flask with a balloon on it is an example of this:		
5. In a chemical reaction t	he arrow points to this:				
Name:	Vaaabulauu				
Period:	v ocabulat y	Closed System Open System Products Reactants	Limiting Reactant Endothermic Exothermic	Law of Conservation of Mass Principle of Definite Proportions	
1. Says that compounds has but not HO or $H_3O$ ).	ave to be only one way $(H_2O is wate)$	er, 6. Says that the science not i	e reacts must equal the pr magic).	oducts (chemistry is	
2. The reactant the is used	l up first and limits the reaction.	7. A reaction the	nat gets hot, like combust	ion (produces energy).	
<ol> <li>A reaction that gets cold (absorbs energy).</li> <li>An open beaker would be an example of this:</li> </ol>				e of this:	
<ul><li>4. The left side of a chemical reaction; the arrow points from here.</li><li>9. A flask with a balloon on it is an example of this:</li></ul>					
5. In a chemical reaction t	he arrow points to this:				
Name:		Closed System	Limiting Reactant	Law of Conservation	
Period:	•	Open System Products Reactants	Endothermic Exothermic	of Mass Principle of Definite Proportions	
1. Says that compounds have to be only one way ( $H_2O$ is water, but not HO or $H_3O$ ).			6. Says that the reacts must equal the products (chemistry is science not magic).		
2. The reactant the is used up first and limits the reaction.		7. A reaction th	7. A reaction that gets hot, like combustion (produces energy).		
3. A reaction that gets colo	d (absorbs energy).	8. An open bea	ker would be an example	e of this:	
4. The left side of a chemic from here.	9. A flask with	9. A flask with a balloon on it is an example of this:			
5. In a chemical reaction t	he arrow points to this:				