

Name: _____

Period: _____

HW—21:AT — After Test
Mr. Murray, IPC
www.aisd.net/smurray

Assigned: Thurs., 11/20/03
Due: Mon., 11/24/03

Type of Reaction	Balance the reactions:
1. _____	____ S ₈ + ____ O ₂ → ____ SO ₃
2. _____	____ C ₃ H ₈ + ____ O ₂ → ____ CO ₂ + ____ H ₂ O
3. _____	____ Na + ____ H ₂ O → ____ Na(OH) + ____ H ₂
4. _____	____ KClO ₃ → ____ KCl + ____ O ₂
5. _____	____ FeCl ₃ + ____ Na(OH) → ____ Fe(OH) ₃ + ____ NaCl

6. You mix vinegar and baking soda. It gets cold—
endothermic or exothermic?

Do Vocabulary on the Back

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**Vocabulary
chapter 20 and 21**

Period: _____

Closed System
Open System
Products
Reactants

Limiting Reactant
Endothermic
Exothermic

Law of Conservation
of Mass
Principle of Definite
Proportions

1. Says that compounds have to be only one way (H_2O is water, but not HO or H_3O).
2. The reactant that is used up first and limits the reaction.
3. A reaction that gets cold (absorbs energy).
4. The left side of a chemical reaction; the arrow points from here.
5. In a chemical reaction the arrow points to this:
6. Says that the reactants must equal the products (chemistry is science not magic).
7. A reaction that gets hot, like combustion (produces energy).
8. An open beaker would be an example of this:
9. A flask with a balloon on it is an example of this:

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