

Name: \_\_\_\_\_

Period: \_\_\_\_\_

**HW4:7 Test Review**  
**Mr. Murray, IPC**

**Assigned: Mon., 11/27 and Tues., 11/28**  
**Due: Wed., 11/29 and Thurs., 11/30**

1. Chemical or Physical change?
- A. \_\_\_\_ Salt is dissolved in water.
  - B. \_\_\_\_ A rock breaks apart due to water freezing in cracks.
  - C. \_\_\_\_ When a metal (like silver or iron) oxidizes (combines with oxygen) and tarnishes (makes a dull cloudy coating) or rusts.
  - D. \_\_\_\_ When wax melts.
  - E. \_\_\_\_ During respiration.
2. Exothermic or Endothermic reaction?
- A. \_\_\_\_ Combine two chemicals and they get cold.
  - B. \_\_\_\_ Photosynthesis
  - C. \_\_\_\_ Combine two chemicals and they get hot.
  - D. \_\_\_\_ Burning paper.

3. Given this reaction:  
 $_____ K + _____ H_2O \rightarrow _____ KOH + _____ H_2$
- A. \_\_\_\_ How many hydrogens are there on the product side (before balanced)?
  - B. What kind of reaction is it?
  - C. Balance the reaction.
4. Find the molecular mass of  $K(NO_3)$
5. How many total Hydrogens are there in  $3Mg(OH)_2$ ?

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Page 2

Type of Reaction	Balance the reactions:
7. _____	_____ $MgO \rightarrow$ _____ $Mg +$ _____ $O_2$
8. _____	_____ $C_2H_6 +$ _____ $O_2 \rightarrow$ _____ $CO_2 +$ _____ $H_2O$
9. _____	_____ $Fe_2O_3 +$ _____ $C \rightarrow$ _____ $Fe +$ _____ $CO$
10. _____	_____ $K_2(PO_4) +$ _____ $MgCl_2 \rightarrow$ _____ $Mg(PO_4) +$ _____ $KCl$
11. _____	_____ $CaO +$ _____ $H_2O \rightarrow$ _____ $Ca(OH)_2$



6. Is the above photosynthesis or respiration?
7. How do you know?
8.  $Be + Cl_2 \rightarrow BeCl_2$  How much chloride was used in this reaction?  
3g + ? g                      24 g
9. Give the balanced formula for Potassium oxide.